Chapter

Significant Role of Perovskite Materials for Degradation of Organic Pollutants

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Abstract

The advancement and the use of visible energy in ecological reparation and photodegradation of organic pollutants are being extensively investigated worldwide. Through the last two decades, great exertions have been dedicated to emerging innocuous, economical, well-organized and photostable photocatalysts for ecofriendly reparation. So far, many photocatalysts mostly based on ternary metal oxides and doped with nonmetals and metals with various systems and structures have been described. Among them, perovskite materials and their analogs (layer-type perovskites) include an emerged as semiconductor-based photocatalysts due to their flexibility and simple synthesis processes. This book chapter precisely concentrates on the overall of related perovskite materials and their associated systems; precisely on the current progress of perovskites that acts as photocatalysts and ecofriendly reparation; explores the synthesis methods and morphologies of perovskite materials; and reveals the significant tasks and outlooks on the investigation of perovskite photocatalytic applications.

Keywords: layered-type perovskite materials, photocatalysis, photodegradation, organic pollutants

1. Introduction

Solar energy is one of the primary sources in the field of green and pure energy that points to the power predicament and climate change task. Solar energy consumption is an ecological reconciliation, and then, the chemical change in solar is presence exhaustive, considered throughout global [1, 2]. In general, solar energy is renewed into a wide range of developments, such as degradation of organic pollutants as photocatalysis, splitting of water molecules for producing clean energy, and reduction of CO_2 gas [3, 4]. Consuming a similar perception, metal-oxide photocatalysis has also been widely examined for possible exertions in ecological restitution as well as the photodegradation and elimination of organic toxins in the aquatic system [5, 6], decrease of bacterial inactivation [7–9], and heavy metal ions [10–12]. Throughout the earlier few years, excellent applications have been dedicated to evolving well-organized, less expensive, and substantial photocatalysts, particularly those that can become active under visible light such as $NaLaTiO_6$, $Ag_3PO_4/BaTiO_3$, $Pt/SrTiO_3$, $SrTiO_3$-TiN, noble-metal-$SrTiO_3$ composites, $GdCoO_3$, orthorhombic perovskites $LnVO_3$ and $Ln_{1-x}Ti_xVO_3$ (Ln = Ce, Pr, and Nd), $Ca_{0.6}Ho_{0.4}MnO_3$, Ce-doped $BaTiO_3$, fluorinated Bi_2WO_6, graphitic

carbon nitride-Bi_2WO_6, $BaZrO_{3-\delta}$, $CaCu_3Ti_4O_{12}$, [13–24], graphene-doped perovskite materials, and nonmetal-doped perovskites [25]. Furthermore, directed to years extended exhaustive investigation exertions on the pursuit of innovative photocatalytic systems, particularly that can produce the overall spectrum of visible-light. Out of a vast assemblage of photocatalysts, perovskite or layered-type perovskite systems and its analogs include a better candidate for capable semiconductor-based photocatalysts due to their framework easiness and versatility, excellent photostability, and systematic photocatalytic nature. In general, the ideal perovskite structure is cubic, and the formula is ABO_3. Where A is different metal cations having charge +1 or +2 or +3 nature and B site occupies with tri or tetra and pentavalent nature, which covers the whole family of perovskite oxide materials by sensibly various metal ions at A and B locations [26], aside from a perfect cubic perovskite system, basic alteration perhaps persuaded by several cations exchange. Such framework alteration could undoubtedly vary the photophysical, optical, and photocatalytic activities of primary oxides.

Moreover, a sequence of layered-type perovskite materials contains many 2D blocks of the ABO_3 framework, which are parted by fixed blocks. The scope of formulating multicomponent perovskite systems by whichever fractional change of cations in A and B or both positions or injecting perovskite oxides into a layered-type framework agrees scientists investigate and control the framework of crystals and the correlated electronic and photocatalytic activities of the perovskites. So far, hundreds of various types of perovskite or perovskite-based catalysts have been published, and more outstandingly, some ABO_3-related materials became renowned with "referred" accomplishment for catalytic activities. Thus, these systems (perovskite materials) have exposed highly capable of upcoming applications on the source of applying more attempts to them. While several outstanding reviews mean that explained that perovskites performed as photocatalyst for degradation of organic pollutants [27–30], only an insufficient of them content consideration to inorganic perovskite (mostly ABO_3-related) photocatalysts [31–33]. A wide range of tagging and complete attention of perovskite materials, for example, layered-type perovskite acting as photocatalysts, is relatively deficient. The purpose of this book chapter is to precise the current progress of perovskite-based photocatalysts for ecological reparation, deliberate current results, and development on perovskite oxides as catalysts, and allow a view on the upcoming investigation of perovskite materials. After a short outline on the wide-ranging structure of perovskite oxides, it was stated that perovskites act as a photocatalyst that are incorporated, arranged and explored based on preparation methods [29, 34], photophysical properties based on bandgap energies, morphology-based framework and the photocatalytic activities depends on either UV or visible light energy of the semiconducting materials. Finally, this chapter is based on the current advancement and expansion of perovskite photocatalytic applications under solar energy consumption. The potential utilization, new tasks, and the research pathway will be accounted for the final part of the chapter [35].

2. Results and discussion

2.1 Details of perovskite oxide materials

2.1.1 Perovskite frameworks

The standard system of perovskite-based materials could be designated as ABO_3, where the A and B are cations with 12-fold coordinated and 6-fold

coordinated to concerning oxygen anions. **Figure 1a** describes the typically coordinated basic of the ABO_3 system, which consists of a 3D system, BO_6 octahedra as located at corner, and at the center, A cation are occupied. Within the ABO_3 system, the A cation usually is group I and II or a lanthanide metal, whereas the B is commonly a transition metal ion. The tolerance factor $(t) = 1$ calculated by using an equation $t = (rA + rO)/\sqrt{2}\,(rB + rO)$, where r_O, r_A, and r_B are the radii of respective ions A and B and oxygen elements for a cubic crystal structure ABO_3 perovskite system [36].

For constituting a stable perovskite, it is typically the range of t value present in between 0.75 and 1.0. The lower value of t builds a somewhat slanted perovskite framework with rhombohedral or orthorhombic symmetry. In the case of t, it is approximately 1; then, perovskite structure is an ideal cubic system at high temperatures. Even though the value of t, obtained by the size of metal ion, is a significant guide for the permanency of perovskite systems, the factor of octahedral (u) $u = r_B/r_O$ and the role of the metal ions composition of A and B atoms and the coordination number of respective metals are considered [37]. Given the account of those manipulating factors and the electro-neutrality, the ABO_3 perovskite can hold a broad variety of sets of A and B by equal or dissimilar oxidation states and ionic radii. Moreover, the replacement of A or B as well as both the cations could be partly by the doping of various elements, to range the ABO_3 perovskite into a wide-ranging family of $A^1_m A^1_{1-m} B^1_n B^1_{1-n} O_{3\pm\delta}$ [38]. The replacement of several cations into the either A or B positions could modify the structure of the original system and therefore improve the photocatalytic activities [23]. After various metal ions in perovskite oxide are doped, the optical and electronic band positions, which influence the high impact on the photocatalytic process, are modified [24].

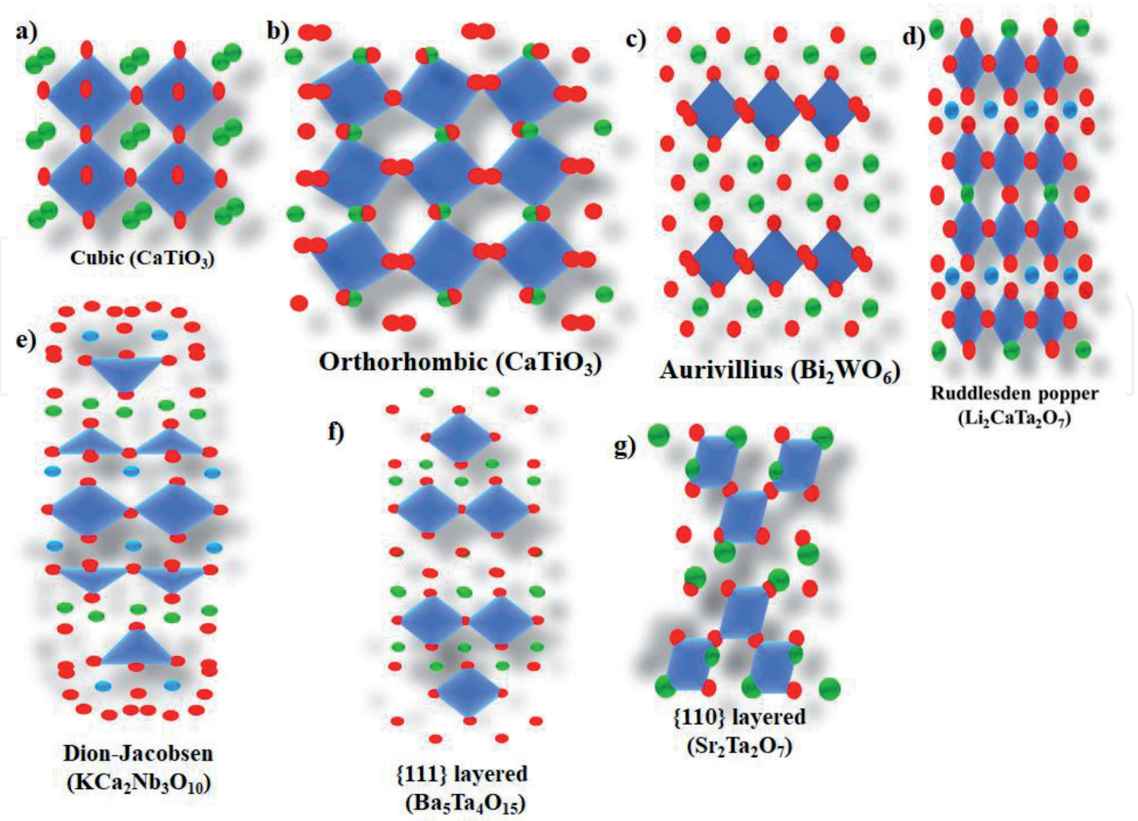

Figure 1.
Both crystal and layered type perovskite oxides (blue small balls: A-site element; dark blue squares: BO_6 octahedra with green and red balls are oxygen).

2.1.2 Layered perovskite-related systems

Moreover, to the overall ABO_3 system, further characteristic polymorphs of the perovskite system are Brownmillerite (BM) ($A_2B_2O_5$) framework [39]. BM is a type of oxygen-deficient perovskite, in which the unit cell is a system of well-organized BO_4 and BO_6 units. The coordination number of cations occupied by A-site was decreased to eight because of the oxygen deficiency. Perovskite (ABO_3) oxides have three dissimilar ionic groups, construction for varied and possibly useful imperfection chemistry. Moreover, the partial replacement of A and B ions is permitted even though conserving the perovskite system and shortages of cations at the A-site or of oxygen anions are common [40]. The Ion-exchange method is used for the replacement of existing metal ions with similar sized or dissimilar oxidation states; then, imperfections can be announced into the system. The imperfection concentrations of perovskites could be led by doping of different cations [24]. Oxygen ion interstitials or vacancies could be formed by the replacement of B-position cations with higher or lower valence, respectively, fabricating new compounds of $AB_{(1-m)}B_m^{I}O_{3-\delta}$ [41]. A typical oxygen-deficient perovskite system is Brownmillerite ($A_2B_2O_5$), in which one part of six oxygen atoms is eliminated. Moreover, the replacement of exciting a site cation to new cation with higher oxidation state metal ions then the formed new materials with new framework with different stoichiometry is $A_{1-m}A_m^{I}BO_3$ [41]. In the case of the replacement of A-site ions with smaller oxidation state cations, consequences in oxygen-deficient materials with new framework such as $A_{1-m}A_m^{I}BO_{3-x}$ are developed. Thermodynamically, the replacement of B-position vacancies in perovskite systems is not preferable due to the compact size and the high charge of B cations [42]. A-position vacancies are more detected due to the BO_3 range in perovskite system forms a stable network [43]; the 12 coordinated sites can be partly absent due to bigger-size A cations. Lately, presenting suitable imperfections on top of the surface of perovskite oxides has been thoroughly examined as a means of varying the bands' position and optical properties of the starting materials. For this reason, perovskite materials afford a tremendous objective for imperfection originating to vary the photocatalytic activity of perovskite material-based photocatalysts [44].

The typical formula for the furthermost recognized layered perovskite materials is designated as $A_{n+1}B_nO_{3n+1}$ or $A_2^{I}A_{n-1}BnO_{3n+1}$ (Ruddlesden-Popper (RP) phase), $A^{I}[A_{n-1}B_nO_{3n+1}]$ (Dion-Jacobson (DJ) phase) for {100} series, ($A_nB_nO_{3n+2}$) for {110} series and ($A_{n+1}B_nO_{3n+3}$) for {111}, and $(Bi_2O_2)(A_{n-1}B_nO_{3n+1})$ (Aurivillius phase) series. In these systems, n represents the number of BO_6 octahedra that duration a layer, which describes the width of the layer. Typical samples of these layered systems are revealed in **Figure 1c–g**. For RP phases, their frameworks consist of $A^{I}O$ as the spacing layer for the intergrowth ABO_3 system. These materials hold fascinating properties such as ferroelectricity, superconductivity, magnetoresistance, and photocatalytic activity. Sr_2SnO_4 and $Li_2CaTa_2O_7$ systems are materials of simple RP kind photocatalysts. A common formula for DJ phase is $A^{I}[A_{n-1}B_nO_{3n+1}]$ (n > 1), where A^{I} splits the perovskite-type slabs and is characteristically a monovalent alkali cation. The typical DJ kind photocatalysts are $RbLnTa_2O_7$ (n = 2) and $KCa_2Nb_3O_{10}$ (n = 3). Associates of the $A_nB_nO_{3n+2}$ and $A_{n+1}B_nO_{3n+3}$ structural sequences with dissimilar layered alignments have also been recognized in some photocatalysts like $Sr_2Ta_2O_7$ and $Sr_5Ta_4O_{15}$ (n = 4). For Aurivillius phases, their frameworks are constructed by one after another fluctuating layers of $[Bi_2O_2]^{2+}$ and virtual perovskite blocks. Bi_2WO_6 and $BiMoO_6$ (n = 1), found as the primary ferroelectric nature for Aurivillius materials, lately have been extensively investigated as visible light photocatalysts.

2.2 Perovskite systems for photocatalysis

A broad array of perovskite photocatalysts have been advanced for organic pollutant degradation in the presence of ultraviolet or visible-light-driven through the last two decades [45]. These typical examples and brief investigational consequences on perovskites are concise giving to their systems, then perovskite materials categorized into six groups. Precisely, ABO_3-type perovskites, AA^IBO_3, A^IABO_3, ABB^IO_3 and $AB(ON)_3$-type perovskites, and $AA^IBB^{II}O_3$-type perovskites are listed in **Table 1**.

2.3 Photocatalytic properties perovskite oxides

$NaTaO_3$ has been a standard perovskite material for a well-organized UV-light photocatalyst for degradation of organic pollutants and production of H_2 and O_2 through water splitting [46–57]. It can be prepared by various methods such as solid-state [46–48, 53, 56], hydrothermal [49, 52, 54, 55], molten salt [57] and sol-gel [50, 51] and with wide bandgap of 4.0 eV. In order to enhance the surface area of $NaTaO_3$ bulk material, many investigators tried to use further synthetic ways to make nanosized particles as an additional study on the $NaTaO_3$ photocatalyst for degradation of organic pollutants. Kondo et al. prepared a colloidal range of $NaTaO_3$ nanoparticles consuming three-dimensional mesoporous carbon as a pattern, which was pretend by the colloidal arrangement of silica nanospheres. After calcining the mesoporous carbon matrix, a colloidal arrangement of $NaTaO_3$ nanoparticles with a range of 20 nm and a surface area of 34 $m^2\ g^{-1}$ was attained. C-doped $NaTaO_3$ material was tested for degradation of NO_x under UV light [36]. Several titanates such as $BaTiO_3$ [58–60], Rh or Fe-doped $BaTiO_3$ [61, 62], $CaTiO_3$ [63, 64] and Cu [65], Rh [66], Ag and La-doped $CaTiO_3$ [67], and $PbTiO_3$ [68, 69] were also described as UV or visible light photocatalysts. Magnetic $BiFeO_3$, recognized as the one of the multi-ferric perovskite materials in magnetoelectric properties, was also examined as a visible light photocatalyst for photodegradation of organic pollutants because of small bandgap energy (2.2 eV) [70–79]. In a previous account, $BiFeO_3$ with a bandgap of around 2.18 eV produced by a citric acid-supported sol-gel technique has revealed its visible-light-driven photocatalytic study by the disintegration of methyl orange dye [70]. The subsequent investigations on $BiFeO_3$ are primarily concentrated on the synthesis of new framework $BiFeO_3$ with various morphologies. For instance, Lin and Nan et al. prepared $BiFeO_3$ unvarying microspheres and microcubes by a using hydrothermal technique as revealed in **Figure 2** [73].

The bandgap energies of $BiFeO_3$ compounds were found to be about 1.82 eV for $BiFeO_3$ microspheres and 2.12–2.27 eV for microcubes. This indicated that the absorption edge was moved toward the longer wavelength that is influenced by the crystal-field strength, particle size, and morphology. The microcube material showed the maximum photocatalytic degradation performance of congo red dye under visible-light irradiation due to the quite low bandgap energy. Further, a simplistic aerosol-spraying method was established for the synthesis of mesoporous $BiFeO_3$ hollow spheres with improved activity for the photodegradation of RhB dye and 4-chlorophenol, because of improved light absorbance ensuing from various light reflections in a hollow chamber and a very high surface area [71]. Moreover, a unusually improved water oxidation property on Au nanoparticle-filled $BiFeO_3$ nanowires under visible-light-driven was described [77]. The Au-$BiFeO_3$ hybrid system was encouraged by the electrostatic contact of negatively charged Au nanoparticles and positively charged $BiFeO_3$ nanowires at pH 6.0 giving to their various isoelectric points. An improved absorbance between 500 and 600 nm was found for Au/$BiFeO_3$ systems because of the characteristic Au surface plasmon band

existing visible light region then which greater influenced in the photodegradation of organic pollutants. Also, the study of photoluminescence supported improvement of the photocatalytic property due to the effective charge transfer from $BiFeO_3$ to Au. Even though Ba, Ca, Mn, and Gd-doped $BiFeO_3$ nanomaterials have exhibited noticeable photocatalytic property for the degradation of dyes [80–84], several nano-based $LaFeO_3$ with various morphologies such as nanoparticles, nanorods, nanotubes, nanosheets, and nanospheres have also been synthesized for visible light photocatalysts for degradation of organic dyes [85–93]. Sodium bismuth titanate ($Bi_{0.5}Na_{0.5}TiO_3$) has been extensively used for ferroelectric and piezoelectric devices. It was also investigated as a UV-light photocatalyst with a bandgap energy of 3.0 eV [94–97]. Hierarchical micro/nanostructured $Bi_{0.5}Na_{0.5}TiO_3$ was produced by in situ self-assembly of $Bi_{0.5}Na_{0.5}TiO_3$ nanocrystals under precise hydrothermal conditions, through the evolution mechanism was examined in aspect means that during which the growth mechanism was studied [95]. It was anticipated that the hierarchical nanostructure was assembled through a method of nucleation and growth and accumulation of nanoparticles and following in situ dissolution-recrystallization of the microsphere type nanoparticles with extended heating period and enhanced temperature or basic settings. The 3D hierarchical $Bi_{0.5}Na_{0.5}TiO_3$ showed very high photocatalytic activity for the decomposition of methyl orange dye because of the adsorption of dye molecules and bigger surface area. The properties of $Bi_{0.5}Na_{0.5}TiO_3$ were also assessed by photocatalytic degradation of nitric oxide in the gas phase [95]. $La_{0.7}Sr_{0.3}MnO_3$, acting as a photocatalyst, was examined for solar light-based photocatalytic decomposition of methyl orange [96–98]. In addition, $La_{0.5}Ca_{0.5}NiO_3$ [99], $La_{0.5}Ca_{0.5}CoO_{3-\delta}$ [100], and $Sr_{1-x}Ba_xSnO_3$ (x = 0–1) [101] nanoparticles were synthesized for revealing improved photocatalytic degradation of dyes. A-site strontium-based perovskites such as $SrTi_{1-x}Fe_xO_{3-\delta}$, $SrTi_{0.1}Fe_{0.9}O_{3-\delta}$, $SrNb_{0.5}Fe_{0.5}O_3$, and $SrCo_{0.5}Fe_{0.5}O_{3-\delta}$ compounds were prepared through solid-state reaction and sol-gel approaches, and were examined for the degradation of organic pollutants under visible light irradiation [102–105]. Also, some other researchers modified A-site with lanthanum-based perovskites such as $LaNi_{1-x}Cu_xO_3$ and $LaFe_{0.5}Ti_{0.5}O_3$ were confirmed as effective visible light photocatalysts for the photodegradation of p-chlorophenol [91, 106, 107]. The other ABB^IO_3 kind photocatalysts with $Ca(TiZr)O_3$ [108], $Ba(ZrSn)O_3$ [109], $Na(BiTa)O_3$ [110], $Na(TiCu)O_3$ [111], $Bi(MgFeTi)O_3$ [112], and $Ag(TaNb)O_3$ [113] have also been studied. Related to AA^IBO_3-type perovskites, the ABB^IO_3 kind system means that BI-site substitution by a different cation is another option for tuning the physicochemical or photocatalytic properties of perovskites materials as photocatalyst, due to typically the B-position cations in ABO_3 mostly regulate the position of the conduction band, moreover to construct the structure of perovskite system with oxygen atoms. The band positions of photocatalyst can be magnificently modified by sensibly coalescing dual or ternary metal cations at the B-position, or changing the ratio of several cations, which has been fine verified by the various materials as mentioned above. More studies on ABB^IO_3 kind of photocatalysts are projected to show their new exhilarating photocatalytic efficiency.

The mesoporous nature of $LaTiO_2N$ of photocatalyst attended due to thermal ammonolysis process of $La_2Ti_2O_7$ precursor from polymer complex obtained from the solid-state reaction. The oxynitride analysis revealed that the pore size and shape, lattice defects and local defects, and oxidation states' local analysis related between morphology and photocatalytic activity were reported by Pokrant et al. [114]. Due to the high capability of accommodating an extensive array of cations and valences at both A- and B-sites, ABO_3-kind perovskite materials are capable materials for fabricating solid-solution photocatalysts. On the other hand, equally the A and B cations can be changed by corresponding cations subsequent in a perovskite with the formula of $(ABO_3)_x(A^IB^IO_3)_{1-x}$. Additional solid solution examples with $CaZrO_3$–$CaTaO_2N$

[115], $SrTiO_3$–$LaTiO_2N$ [116], $La_{0.8}Ba_{0.2}Fe_{0.9}Mn_{0.1}O_{3-x}$ [117], $Na_{1-x}La_xFe_{1-x}Ta_xO_3$ [118], $Na_{0.5}La_{0.5}TiO_3$–$LaCrO_3$ [119], Cu-$(Sr_{1-y}Na_y)$-$(Ti_{1-x}Mo_x)O_3$ [120], $Na_{1-x}La_xTa_{1-x}Cr_xO_3$ [121], $BiFeO_3$–$(Na_{0.5}Bi_{0.5})TiO_3$ [122], and $Sr_{1-x}Bi_xTi_{1-x}Cr_xO_3$ [123] have been used as photocatalysts for splitting of water molecules under visible light.

2.4 Photocatalytic activity of layered perovskite materials

In the general formula of the RP phase, $A_{n-1}A_2^IB_nO_{3n+1}$, A and A^I are alkali, alkaline earth, or rare earth metals, respectively, while B states to transition metals. A and A^I cations are placed in the perovskite layer and boundary with 12-fold cuboctahedral and 9-fold coordination to the anions, respectively, whereas B cations are sited inside the perovskite system with anionic squares, octahedra, and pyramids. The tantalum-based RP phase materials have been examined as photocatalysts for degradation of organic pollutants under UV light irradiation conditions; such materials are $K_2Sr_{1.5}Ta_3O_{10}$ [124], $Li_2CaTa_2O_7$ [125], $H_{1.81}Sr_{0.81}Bi_{0.19}Ta_2O_7$ [126], and N-alkyl chain inserted $H_2CaTa_2O_7$ [127]. A series of various metals and N-doped perovskite materials were synthesized, such as Sn, Cr, Zn, V, Fe, Ni, W, and N-doped $K_2La_2Ti_3O_{10}$, for photocatalysis studies under UV and visible light irradiation [128–133]. Still, only Sn-doping efficiently decreased the bandgap energy of $K_2La_2Ti_3O_{10}$ from 3.6 eV to 2.7 eV. The bandgap energy of N-doped $K_2La_2Ti_3O_{10}$ was measured to be around 3.4 eV. Additional RP phase kind titanates like Sr_2SnO_4 [134], $Sr_3Ti_2O_7$ [135], Cr-doped Sr_2TiO_4 [136], $Sr_4Ti_3O_{10}$ [137], $Na_2Ca_2Nb_4O_{13}$ [138], and Rh- and Ln-doped $Ca_3Ti_2O_7$ [139] have also been examined. Bi_2WO_6 (2.8 eV) shows very high oxygen evolution efficacy than Bi_2MoO_6 (3.0 eV) from aqueous $AgNO_3$ solution under visible-light-driven. Because of the appropriate bandgap energy, comparatively elevated photocatalytic performance, and good constancy, Bi_2MO_6 materials have been thoroughly examined as the Aurivillius phase kind that acts as photocatalysts under visible light. In this connection, hundreds of publications associated to the Bi_2MoO_6 and Bi_2WO_6 act as photocatalysts so far reported. Most of the investigations in the reports are concentrated on the synthesis of various nanostructured Bi_2MoO_6 and Bi_2WO_6 as well as nanofibers, nanosheets, ordered arrays, hollow spheres, hierarchical architectures, inverse opals, and nanoplates, etc., by various synthesis techniques like solvothermal, hydrothermal, electrospinning, molten salt, thermal evaporation deposition, and microwave. All these methods of hydrothermal process have been frequently working for the controlled sizes, shapes, and morphologies of the particles. The photocatalytic properties of these perovskite materials are mostly examined by the photodegradation of organic pollutants. Moreover, the investigations on the simple Bi_2MoO_6 and Bi_2WO_6, doped with various metals and nonmetals such as Zn, Er, Mo, Zr, Gd, W, F, and N, into Bi_2MoO_6 and Bi_2WO_6 was studied for increasing the photocatalytic performance under visible light. Therefore, these Bi_2MO_6-based photocatalysts is not specified here, due to further full deliberations that can be shown in many reviews [140–142].

$ABi_2Nb_2O_9$ where A is Ca, Sr, Ba and Pb is other type of the AL-like layered perovskite material [143–150]. The bandgap energy of $PbBi_2Nb_2O_9$ is 2.88 eV and originally described as an undoped with single-phase layered-type perovskite material used as photocatalyst employed under visible light irradiation [144]. $Bi_5FeTi_3O_{15}$ is also Aurivillius (AL) type multi-layered nanostructured perovskite material with a low bandgap energy (2.1 eV) and also shows photocatalytic activity under visible light [151, 152]. Mostly, these materials were synthesized using the hydrothermal method that has been frequently working for the controlled shapes such as flower-like hierarchical morphology, nanoplate-based, and the complete advance process from nanonet-based to nanoplate-based micro-flowers was shown. The photocatalytic activity of $Bi_5FeTi_3O_{15}$ was studied by the degradation of rhodamine

B and acetaldehyde under visible light [151]. The La substituted $Bi_{5-x}La_xTi_3FeO_{15}$ (x = 1, 2) Al-type layered materials were synthesized through hydrothermal method and these materials were used for photodegradation of rhodamine B under solar-light irradiation [153]. Among all AL-type perovskite materials, only $PbBi_2Nb_2O_9$, Bi_2MO_6 (M = W or Mo), and $Bi_5Ti_3FeO_{15}$ are very high photocatalytic active under visible-light-driven due to low bandgap energy and photostability. Another type of layered perovskite material is Dion-Jacobson phase (DJ), a simple example is $CsBa_2M_3O_{10}$ (M = Ta, Nb) and oxynitride crystals used for degradation of caffeine from wastewater under UVA- and visible-light-driven [154]. Similarly, another DJ phase material such means Dion–Jacobsen (DJ) as $CsM_2Nb_3O_{10}$ (M = Ba and Sr) and also doped with nitrogen used for photocatalysts for degradation of methylene blue [155]. Zhu et al. prepared tantalum-based {111}-layered type of perovskite material such as $Ba_5Ta_4O_{15}$ from hydrothermal method, which has been frequently employed for the controlled shape like hexagonal structure with nanosheets and used as photocatalyst for photodegradation of rhodamine B and gaseous formaldehyde [156]. Pola et al. synthesized a layered-type perovskite material constructed on $A^IA^{II}Ti_2O_6$ (A^I = Na or Ag or Cu and A^{II} = La) structure for the photodegradation of several organic pollutants and industrial wastewater under visible-light-driven [157–162].

Perovskite system	Synthesis process	Light source	Pollutants	References
$NaTaO_3$	HT	UV	CH_3CHO	[163]
La-doped $NaTaO_3$	SG	UV	MB	[164]
La-doped $NaTaO_3$	HT	UV	MB	[165]
Cr-doped $NaTaO_3$	HT	UV	MB	[166]
Eu-doped $NaTaO_3$	SS	UV	MB	[167]
Bi-doped $NaTaO_3$	SS	UV	MB	[168]
N-doped $NaTaO_3$	SS	UV	MB	[169]
C-doped $NaTaO_3$	HT	Visible	NO_x	[36]
N/F co-doped $NaTaO_3$	HT	UV	RhB	[170]
$SrTiO_3$	HT	UV	RhB	[42, 43, 171]
Fe-doped $SrTiO_3$	SG	Visible	RhB	[172]
N-doped $SrTiO_3$	HT	Visible	MB, RhB, MO	[173]
F-doped $SrTiO_3$	BM	Visible	NO	[174]
Ni/La-doped $SrTiO_3$	SG	Visible	MG	[175]
S/C co-doped $SrTiO_3$	SS	Visible	2-Propanol	[176]
N/La-doped $SrTiO_3$	SG	Visible	2-Propanol	[177]
Fe-doped $SrTiO_3$	ST	Visible light	TC	[178]
$SrTiO_3/Fe_2O_3$	HT	Visible	TC	[179]
$BaTiO_3$	SG	UV	Pesticide	[36]
$BaTiO_3$	SG	UV	Aromatics	[58]
$BaTiO_3$	HT	UV	MO	[58]
$KNbO_3$	HT	Visible	RhB	[180]
$KNbO_3$	HT	UV	RhB	[181]
$KNbO_3$	HT	Visible	MB	[182]
$NaNbO_3$	SS	UV	RhB	[183]
$NaNbO_3$	Imp.	UV	2-Propanol	[184]
$NaNbO_3$	SS	UV	MB	[185]

Perovskite system	Synthesis process	Light source	Pollutants	References
N-doped $NaNbO_3$	SS	UV	2-Propanol	[186–188]
Ru-doped $NaNbO_3$	HT	Visible	Phenol	[189]
$AgNbO_3$	SS	UV	MB	[190]
La-doped $AgNbO_3$	SS	Visible	2-Propanol	[191]
$BiFeO_3$	SG	UV-Vis	MO, RhB, 4-CP	[69– 77]
Ba-doped $BiFeO_3$	ES	Visible	CR	[79]
Ca-doped $BiFeO_3$	ES	Visible	CR	[80]
Ba or Mn-doped $BiFeO_3$	ES	Visible	CR	[82]
Ca or Mn-doped $BiFeO_3$	HT	UV-Visible	RhB	[82]
Gd-doped $BiFeO_3$	SG	Visible	RhB	[83]
$LaFeO_3$	Comb.	UV	Methyl phenol	[84]
$LaFeO_3$	SG	Visible	RhB	[85]
$LaFeO_3$	HT	Visible	RhB, MB, chlorophenol	[86, 90, 91, 192]
Ca-doped $LaFeO_3$	SS	Visible	MB	[92]
$LnFeO_3$ (Pr,Y)	SG	Visible	RhB	[193]
$SrFeO_{3-x}$	US	Visible	Phenol	[194]
$SrFeO_3$	SS	Visible	MB	[195]
$BaZrO_3$	SG	UV	MB	[196]
$BaZrO_3$	HT	UV	MO	[197]
$ATiO_3$ (A = Fe, Pb) and $AFeO_3$ (A = Bi, La, Y)	SG	Visible	MB	[198]
$Zn_{0.9}Mg_{0.1}TiO_3$	SG	Visible	MB	[199]
$SrTiO_3$ nanocube-coated CdS microspheres	HT	Visible	Antibiotic pollutants	[200]
Ag/AgCl/$CaTiO_3$	HT	Visible	RhB	[201]
TiO_2-coupled $NiTiO_3$	SS	Visible	MB	[202]
$ZnTiO_3$	HT	UV	MO and PCP	[203]
Mg-doped $BaZrO_3$	SS	UV	MB	[204]
$SrSnO_3$	MW	UV	MO	[205]
$LaCoO_3$	MW	Visible	MO	[206]
$LaCoO_3$	Ads.	UV	MB, MO	[207]
$LaCoO_3$	ES	UV	RhB	[208]
$LaNiO_3$	SG	Visible	MO	[209]
$Bi_{0.5}Na_{0.5}TiO_3$	HT	UV	MO	[93]
$La_{0.7}Sr_{0.3}MnO_3$	SG	Solar light	MO	[97]
$La_{0.5}Ca_{0.5}NiO_3$	SG	UV	RB5	[98]
$La_{0.5}Ca_{0.5}CoO_3$	SG	UV	CR	[99]
$Sr_{1-x}Ba_xSnO_3$	SS	UV	Azo-dye	[100]
$BaCo_{1/2}Nb_{1/2}O_3$	SG	Visible	MB	[210]
$Ba(In_{1/3}Pb_{1/3}M_{1/3})O_3$ (M = Nb and Ta)	SS	Visible	MB, 4-CP	[211]
$A(In_{1/3}Nb_{1/3}B_{1/3})O_3$ (A = Sr, Ba; B = Sn, Pb)	SS	Visible	MB, 4-CP	[212]
$SrTi_{1-x}Fe_xO_{3-\delta}$	SS	Visible	MB	[102]

Perovskite system	Synthesis process	Light source	Pollutants	References
$SrTi_{0.1}Fe_{0.9}O_{3-\delta}$	SG	Solar light	MO	[103]
$SrFe_{0.5}Co_{0.5}O_{3-\delta}$	SG	Solar light	CR	[213]
$LaFe_{0.5}Ti_{0.5}O_3$	SG	UV	Phenol	[90]
$Bi(Mg_{3/8}Fe_{2/8}Ti_{3/8})O_3$	MS	Visible	MO	[110]
$LaTi(ON)_3$	SG	Visible	Acetone	[214]
$(Ag_{0.75}Sr_{0.25})(Nb_{0.75}Ti_{0.25})O_3$	SS	Visible	CH_3CHO	[215]
$La_{0.8}Ba_{0.2}Fe_{0.9}Mn_{0.1}O_{3-x}$	SG	Solar light	MO	[115]
$Cu\text{-}(Sr_{1-y}Na_y)(Ti_{1-x}Mo_x)O_3$	HT	Visible	Propanol	[118]
$BiFeO_3–(Na_{0.5}Bi_{0.5})TiO_3$	SG	Visible	RhB	[120]
$SrBi_2Nb_2O_9$	SG SS	UV	Aniline, RhB	[145, 146]
$ABi_2Nb_2O_9$ (A = Sr, Ba)	SG	UV	MO	[147]
$Bi_5Ti_3FeO_{15}$	HT SS	Visible	RhB, CH_3CHO IPA	[149, 150]
$Bi_{5-x}La_xTi_3FeO_{15}$	SS	Solar light	RhB	[151]
$Bi_3SrTi_2TaO_{12}$ $Bi_2LaSrTi_2TaO_{12}$	SS	UV	RhB	[216]
$Ba_5Ta_4O_{15}$	HT	UV	RhB	[154]
N-doped $Ln_2Ti_2O_7$ (Ln = La, Pr, Nd)	HT	Visible	MO	[217]
$CdS/Ag/Bi_2MoO_6$	SG	Visible	RhB	[218]

SS: solid state; HT: hydrothermal; SG: sol-gel; BM: ball-milling; ES: electronspun; MW: microwave; Comb.: combustion; US: ultrasonic; MS: molten salt; Imp.: impregnation; Ads.: adsorption; ST: solvothermal; RhB: rhodamine B; MO: methyl orange; MB: methylene blue; 4-cp: 4-chlorophenol; MG: malachite green; CR: congo red; NO: nitrogen monoxide; PA: isopropyl alcohol; TC: tetracycline; and PCP: pentachlorophenol.

Table 1.
Perovskite materials used as photocatalysts (ABO_3, AA^IBO_3, AA^IBO_3, ABB^IO_3, $AB(ON)_3$, and $AA^IBB^{II}O_3$) for degradation of pollutants.

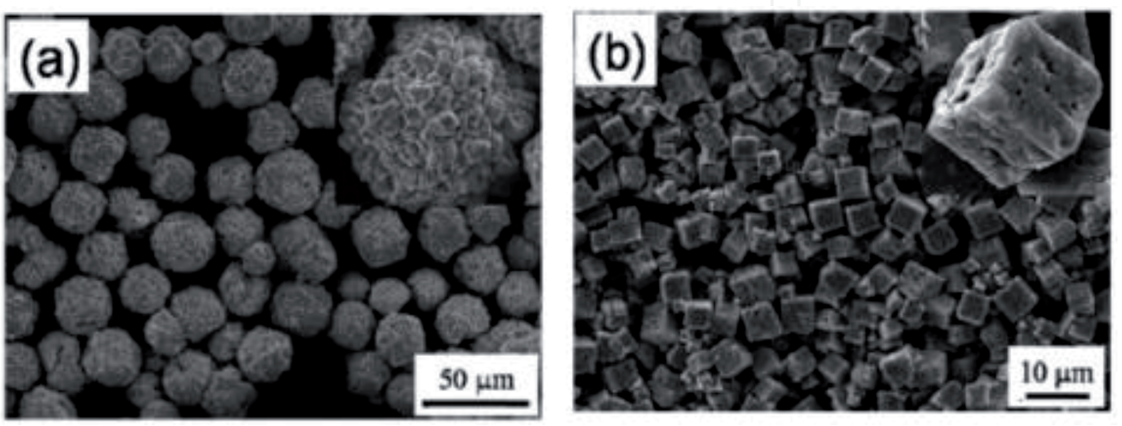

Figure 2.
SEM patterns of $BiFeO_3$: (a) microspheres and (b) microcubes. The intensified pictures are revealed in the upper part inserts. Recopied with consent from Ref. [147]. Copyright © 2010, American Chemical Society.

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